

When analyzing allylic halogenation reactions (NBS and $h\nu$)

1. Consider all possible allylic radicals that can be formed.
2. Analyze all contributing structures for all of the allylic radicals.
3. Add a Br atom at the site of the unpaired electron for all contributing structures for all of the allylic radicals.
4. From all of the possible products, the predominant product is the one **THAT IS THE MOST STABLE ALKENE** – the most substituted alkene – alkyl groups stabilize alkenes – *trans* over *cis*.
5. Note: It is OK if the product you choose derives from an allylic radical contributing structure that is a minor contributor. **FOR THIS REACTION WE ONLY CARE ABOUT THE RELATIVE STABILITIES OF THE PRODUCT ALKENES.**

