HIV-1 protease: mechanism and drug discovery

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1 Introduction
It has now been two decades since acquired immunodeficiency syndrome (AIDS) was first reported by the US Center for Diseases Control (CDC). A few years later, it was found that a retrovirus called human immunodeficiency virus (HIV) is the causative agent in AIDS.1 In a short time, AIDS increased to epidemic proportions throughout the world, affecting more than 40 million people today and killing so far more than 22 million (UNAIDS, 2001). Since the outbreak of the AIDS epidemic, tremendous efforts have been directed towards development of antiretroviral therapies that target HIV type 1 in particular (HIV-1). The identification of the HIV retrovirus and the accumulated knowledge about the role of the different elements in its life cycle led researchers around the world to develop inhibitors that target different steps in the life cycle of the virus. One of these targets is HIV-1 protease (HIV PR), an essential enzyme needed in the proper assembly and maturation of infectious virions. Understanding the chemical mechanism of this enzyme has been a basic requirement in the development of efficient inhibitors. In this review, we summarize studies conducted in the last two decades on the mechanism of HIV PR and the impact of their conclusions on the drug discovery processes.

2 The life cycle of HIV
HIV belongs to the class of viruses called retroviruses, which carry genetic information in the form of RNA. HIV infects T cells that carry the CD4 antigen on their surface. The infection of the virus requires fusion of the viral and cellular membranes, a process that is mediated by the viral envelope glycoprotein (gp120, gp41) and receptors (CD4 and coreceptors, such as CCR5 or CXCR4) on the target cell. As the virus enters a cell, its RNA is reverse-transcribed to DNA by a virally encoded enzyme, the reverse transcriptase (RT). The viral DNA enters the cell nucleus, where it is integrated into the genetic material of the cell by a second virally encoded enzyme, the integrase. Activation of the host cell results in the transcription of the viral DNA into messenger RNA, which is then translated into protein. HIV protease, the third virally encoded enzyme, is required in this step to cleave a viral polyprotein precursor into individual mature proteins. The viral RNA and viral proteins assemble at the cell surface into new virions, which then bud from the cell and are released to infect another cell. The extensive cell damage from the destruction of the host’s genetic system to the budding and release of virions leads to the death of the infected cells.

3 HIV protease
3.1 HIV protease: a logical target for AIDS therapy
Unless the HIV life cycle is interrupted by specific treatment, the virus infection spreads rapidly throughout the body, which results in the weakness and destruction of the body’s immune system. From the analysis of the HIV life cycle, one could conclude that there are several steps that might be interfered with, thus stopping the replication of the virus. For example, there are several commercially available drugs that inhibit the enzyme reverse transcriptase (RT). The first class of RT inhibitors is the nucleoside analogs such as AZT, ddI, ddC and d4T. These dideoxy compounds lack the 3'-hydroxy, causing DNA chain termination when they are incorporated into the growing DNA strand. The second class of inhibitors is the non-nucleoside inhibitors (NNIs); these inhibitors are known to bind in a pocket away from the polymerase active site, and are believed to cause a conformational change of the enzyme active site, and thus inhibit its action. Currently, there are three available non-nucleoside reverse transcriptase inhibitors (nevirapine, delavirdine, and efavirenz) for the treatment of AIDS.

Another critical step in the life cycle of HIV is the proteolytic cleavage of the polypeptide precursors into mature enzymes and structural proteins catalyzed by HIV PR. It has been shown that budded immature viral particles that contain catalytically inactive protease cannot undergo maturation to an infective form.2 The necessity of this enzyme in the virus life cycle makes it a promising target for therapy of the HIV infection.3

3.2 Structure of HIV protease
Navia et al. from Merck laboratories were the first group to obtain a crystal structure of HIV PR;4 a more accurate structure was reported subsequently by Kent and coworker.3 HIV PR is a 99 amino acid aspartyl protease which functions as a homodimer with only one active site which is C2-symmetric in the free form. More than 140 structures of the HIV-1 PR, its mutants and enzymes complexed with various inhibitors have been reported so far. A database dedicated to providing structural information about HIV PR has been created at the National Cancer Institute (http://www-fbsc.ncifcrf.gov/HIVdb/). The enzyme homodimer complexed with TL-3 is shown in Fig. 1 (PDB ID: 3TLH). Each monomer contains an extended β-sheet region (a glycine-rich loop) known as the flap, that constitutes in part the substrate-binding site and plays an important role in substrate binding, and one of the two essential asparagine residues, Asp-25 and Asp-25' which lie on the bottom of the cavity. The substrate binds in its extended conformation, in which its interactions with the different amino acids in the flap are optimal.
acid side chains determine the specificity of the enzyme. Using standard nomenclature (Fig. 2), the S1 and S1’ (S2 and S2’, etc.) subsites are structurally equivalent. The two S1 subsites are very hydrophobic, the S2 subsites are mostly hydrophobic except Asp-29, Asp-29’, Asp-30 and Asp-30’. The S3 subsites are adjacent to S1 subsites and are also mostly hydrophobic.

3.3 Mechanism of the HIV protease

Proteases are known to play essential roles in many biological processes. They catalyze the hydrolysis of peptide bonds with high sequence selectivity and catalytic proficiency. These enzymes accomplish their catalysis by two different mechanisms, thus dividing them mechanistically into two broad classes of protease enzymes. The first class of enzymes uses an activated water molecule to attack the amide bond carbonyl of the substrate’s scissile bond. The activation of the water molecule can be achieved either by a zinc cation (the zinc metalloproteinases) or by the two aspartyl β-carboxy groups at the active site (the aspartate proteases). In the second class of proteases, a nucleophlic atom of an amino acid side chain is used to initiate amide hydrolysis. In the first step the nucleophile, which can be a hydroxy group or a thiol, is activated by another amino acid side chain. The activated nucleophile attacks the carbonyl of the scissile amide bond to form an ester or a thioester acyl intermediate. Eventually, this acyl enzyme intermediate is hydrolyzed by a water molecule to the corresponding hydrolysis products. According to several studies, HIV PR in general has been shown to belong to the class of the aspartic proteases. Examining the sequence homology of HIV PR to other cellular aspartic proteases shows that this enzyme has the sequence Asp-Thr-Gly, which is conserved among the aspartic protease enzymes as well as a similar mechanism. Indeed, mutational analysis by several groups of the highly conserved Asp 25 has shown that substituting this residue with Asn, Thr, or Ala, leads to a protein without any proteolytic activity. More support for HIV PR being a member of the aspartic protease family came from the in vitro inhibition of this enzyme by pepstatin, a natural product that selectively inhibits members of this family. Finally, the three-dimensional structure of this enzyme also supported the classification of HIV PR in the aspartic protease family. The dimeric structure, in which each monomer contributes one Asp-Thr-Gly triad to the pseudo-symmetric active site, shows an active site that is indistinguishable from those of the monomeric aspartic proteases.

The catalytic mechanism of the nonviral aspartic proteases has been extensively studied by kinetic methods, affinity labeling and X-ray crystallography. Several mechanisms of action for this family have been proposed, including a mechanism that involves the formation of a covalent intermediate. However, the results from most of these studies are consistent with a general acid-base mechanism, in which the two active site aspartate residues play an essential general acid-base role to activate the water molecule that acts as a nucleophile and attacks the carbonyl carbon of the scissile bond. It is generally believed that this water molecule is located between the active site aspartates, although some have suggested a different nucleophilic water molecule.

The most widely accepted mechanism for aspartic protease has been described by Suguna et al. (Fig. 3). The proposed mechanism is based on the crystal structure of the Rhizopus oryzae aspartic protease complexed with a reduced peptide inhibitor. The pH-rate profile of this enzyme implies that only one of the two active site aspartic acids is unprotonated in the active pH range. In the proposed mechanism the Asp group that is closer to the nucleophilic water molecule was assigned the negative charge (Fig. 3). The nucleophilic water molecule held between the catalytic aspartates, after its activation by the negative aspartate side chain, attacks the carbonyl group in the substrate scissile bond to generate an oxyanion tetrahedral intermediate. Protonation of the scissile amide N atom and rearrangement result in the breakdown of the tetrahedral intermediate to the hydrolysis products. The general acid–base mechanism of the aspartic protease family precludes the use of a Lewis acid such as Zn(II) (as in the case of zinc metalloproteinases) and the formation of covalent acyl intermediates.

Although the HIV PR mechanism shares many features with the rest of the aspartic protease family, the full detailed mechanism of this enzyme remains not fully understood. Jaskolski et al. have proposed a new model of the enzymatic mechanism for the HIV PR enzyme based on the crystal structure of a complex between a chemically synthesized HIV PR and an octapeptide inhibitor U-85548e (H-Val-Ser-Gln-Asn-Leu-Gly-Ile-Ile-Val-OH), as well as other crystal structures of HIV PR complexed with different inhibitors. In this mechanism (Fig. 4), the hydrolysis reaction is viewed as a one-step process during which the nucleophilic water molecule and the acidic proton attack the scissile peptide bond in a concerted manner. The issue of the simultaneous attack from the nucleophile and electrophile is the major difference between this mechanism and other previously proposed mechanisms.

As in the case of all the aspartic protease family, the possibility of covalent catalysis (e.g. Asp-25 attacks directly the carbonyl amide bond) in the chemical mechanism of HIV PR is discouraged. Indeed, Hyland et al. have provided evidence against the formation of a covalent intermediate by studying the incorporation into the transpeptidation product of the reaction of AcSQNYPFLD-GNH₂ from AcSQNYPV-NH₂ and FLDG-NH₂ carried out in H₂O enriched water and catalyzed by HIV PR. Since the incorporation of O₁₈O into the substrate AcSQNYPV-NH₂ (Fig. 5) could be observed under the condition where the re-formation of substrate after hydrolysis was negligible, the involvement of an acyl enzyme intermediate could be excluded. From these studies, it was also found that the HIV PR-catalyzed exchange of O₁₈O from H₂O into the re-formed substrates occurs at a rate 0.01–0.012 times that of the hydrolysis rate. These results are in agreement with the formation of a kinetically competent enzyme-bound amide hydrate intermediate, the collapse of which is the rate-limiting chemical step in the reaction.

We have described a new pyrrolidine α-keto amide core structure as a mechanism-based inhibitor for HIV PR. The crystal structure shows that this inhibitor is bound to HIV PR in its hydrated form (Fig. 6). A ¹³C NMR study in DMSO-d₆-D₂O (5 : 1) has determined that, in the presence of water, the ketone group of this inhibitor remains unhydrated even after 24 h. These results suggest that the hydration takes place within the enzyme-binding site, which means that the catalytic water is incorporated into the inhibitor. These results support the role of the nucleophilic character of the catalytic water molecule in
the reaction pathway and may imply an electrophilic assistance for hydration of carbonyl with low electrophilic character.

There are four possible protonation states of the two catalytic aspartates: the dianionic form (−1, −1), the two monoanionic forms (−1, 0), (0, −1) in which one of the catalytic aspartate groups is anionic and the other protonated, and the diprotonated or neutral form (0, 0). It has been suggested that the protonation state of the two aspartic groups depends on the local environment near the aspartate and is different for different inhibitors. However, it has been suggested also that the difference in pK$_a$s of these residues is more a function of their proximity to each other than of their differing environments.

Aspartic proteases are unique in that they function physiologically over a wide pH range (2–7.4). The pH–rate profile of a model substrate for HIV PR studied by Meek and coworker shows that the two aspartate groups have different pK$_a$ values of 3.1 and 5.2. In contrast to these results, Smith et al. have found in their NMR experiments on the 13C-enriched enzyme in the absence of inhibitor, that the two Asp side chains are equivalent and are both deprotonated at pH 6. However, in the presence of pepstatin inhibitor only one Asp side chain is deprotonated in the pH range 2.5–6.5. Wang et al. have also examined the protonation state of the two aspartic groups, however using the asymmetric inhibitor KNI-272. They also found one catalytic Asp to be protonated and the other unprotonated. In contrast to these results, Yamazaki et al. have studied the ionization state of the two aspartic groups in HIV PR complexed with the symmetric inhibitor DM323. This non-peptidic inhibitor contains two hydroxy groups and forms a completely symmetric complex with the enzyme, in which the side chains of the two aspartates were found to be magnetically equivalent and both protonated over the pH range 2–7. Using ab initio molecular dynamics studies of the pepstatin A–HIV PR complex, Piana et al. have proposed recently that both aspartic groups are protonated, in contrast to what had been reported by Smith et al. An answer to the question of where the acidic proton is located in the free enzyme could not be easily provided since proton locations are generally not resolvable by X-ray crystallography methods.

Meek et al. have shown that the known aspartic protease inactivator 1,2-epoxy-3-(4-nitrophenoxy)propane produced irreversible, time-dependent inactivation of HIV PR, through covalent modification of the enzyme’s aspartyl residue. The pH-dependent kinetics of this inhibitor–enzyme interaction was consistent with having one protonated aspartic group in the active site of the enzyme. Following this observation, Lee and coworkers have designed several inhibitors based on a peptide isostere containing cis-epoxide for the irreversible inactivation of HIV (Fig. 7). These results are surprising, since one would expect the catalytic water molecule to be involved in the epoxide ring opening instead of direct attack by the aspartyl group. Perhaps the alignment of the inhibitor in the active site does not

**Fig. 3** Proposed catalytic mechanism for aspartic proteases.

**Fig. 4** Proposed concerted catalytic mechanism for HIV.
allow enough space between the epoxide ring and the two aspartyl residues to accommodate a water molecule.

Hyland et al. have proposed a detailed chemical mechanism for HIV PR based on kinetic data obtained from solvent kinetic isotope effects, pH–rate and $^{18}$O incorporation studies, combined with known previous structural data of HIV PR (Fig. 8). In this mechanism, Asp-25' exists in the unprotonated state ($pK_a$ 3.1) upon binding to substrate, while the proton on Asp-25 ($pK_a$ 5.2) is positioned to hydrogen bond to the carbonyl oxygen of the substrate, and at the same time, the lytic water is positioned closer to the $\beta$-carboxylic acid of Asp-25'. This mechanism has many similar features in common with the general acid–base mechanism of aspartic proteases reported by other groups. It also resembles in part the concerted mechanism proposed by Jaskólski, in that the amine product is protonated by an active site aspartyl residue. However in
Hyland’s proposal, a discrete enzyme intermediate is formed while, in the concerted mechanism, there is no such inter-
mediate.

Pearl has proposed a mechanism of aspartic proteases similar to what Suga et al. have suggested, however pointing out
the effect of the distortion of the scissile bond on the catalysis.
In this mechanism the distortion of the scissile bond reduces
the double bond character in the C–N bond and polarizes the
carbonyl, rendering it more electrophilic towards the incom-
ing nucleophilic water. The distortion of the amide bond is
stabilized by interactions of the substrate with the extended
binding cleft, mostly by hydrogen bond interactions between
the interior side of the flap and the substrate carboxyls on
either side of the scissile bond.

In HIV PR there are two flaps in the active dimer compared
to the single flap in non-viral aspartic protease. Moreover,
while in the pepsin-like protease there is a direct H-bonding
between the flap and the substrate, in HIV PR the carbonyl
group between P1 and P2 and the carbonyl between P1’ and P2’
both make hydrogen bonds to the same water molecule (Fig. 9).
This water molecule makes another two hydrogen bonds to the
flaps to create approximately tetrahedral geometry. The pres-
ence of this water molecule has been observed in most of the
crystal structures of HIV PR bound to different inhibitors.

It has been proposed that this water molecule is relevant for
catalysis and a key feature in the differences of the mechan-
isms of the retroviral HIV PR and the corresponding cell-
encoded enzymes. The hydrogen bonds of this water molecule
apply strain on the scissile amide bond, causing it to rotate out
of the plane and lose its double bond character, which enhances
its vulnerability towards hydrolysis. Using total chemical syn-
thesis of proteins, HIV PR was synthesized in which the Gly–
Hle N(H)-atom (Fig. 9) was specifically replaced by an
O-atom, thus deleting one of the hydrogen bonds from one of
the flaps to the water molecule. The resulting enzyme with the
single flap–substrate hydrogen bonds was fully active. Based
on these results, it has been suggested that this enzyme may
make use of only one flap in the catalysis, showing a similarity
to pepsin-like enzymes.

4 Mechanism based strategy for drug design
The knowledge accumulated through the last two decades
about the structure and the mechanism of HIV PR has paved
the way towards the development of effective drugs for this
enzyme. In addition, there is a huge amount of information
regarding the design of inhibitors for other aspartic proteases,
The hydroxy group is worth more than 4 kcal mol\(^{-1}\) than 2000 inhibitors have been developed and most of the tetrahedral transition state of the proteolytic reaction. More replacement at the scissile bond that is believed to mimic the cleavage site. The structure of this inhibitor was quite interesting since it had the \(R\) stereochemistry at the carbon bearing the central hydroxy group, which was unique among aspartic protease inhibitors that were known at that time. Both Amprenavir \((K_i = 0.6 \text{ nM})\), which was discovered by Vertex Pharmaceuticals and approved by the FDA in 1999, and Nelfinavir \((K_i = 2 \text{ nM})\), which was discovered by a collaborative effort between Lilly and Agouron and approved by the FDA in 1997, have similar hydroxyethylamine isostere replacement to that found in Saquinavir.

Merck group discovered a novel variation of the hydroxyethylene transition state analog in their new drug, Indinavir \((K_i = 0.56 \text{ nM})\), which was approved by the FDA in 1996. Abbott Laboratories have used the \((S,S,S)\)-aminoalcohol to prepare two of the commercially available drugs: Ritonavir \((K_i = 0.01 \text{ nM})\), which was approved by the FDA in 1996, and Lopinavir \((K_i = 0.003 \text{ nM})\), which is the latest approved HIV PR inhibitor (September 2000). Lopinavir is contained in a protease inhibitor formulation (Klaretra) which includes Ritonavir. Ritonavir is known to inhibit cytochrome P-450 3A, the enzyme responsible for the metabolism of Lopinavir, therefore this combination allows for increased plasma levels of Lopinavir.

The Dupont Merck group has designed a new generation of HIV PR inhibitors based on cyclic urea cores (Fig. 12). The uniqueness of this class of inhibitors is the cyclic urea carboxyl oxygen that mimics the hydrogen bonding features of the key structural water molecule discussed earlier in this review (Fig. 9). The group reasoned that incorporating the cyclic urea core into an inhibitor would replace the flap water molecule, which could lead to better binding energy due to the positive entropic effect that should be provided by the cyclic urea core. These inhibitors also contained the diol functionality as a transition state mimic to interact with the catalytic aspartates. Several inhibitors based on these concepts were developed and found to be potent against HIV PR. Although some of these inhibitors were clinically tested, none of them have as yet reached the market.

### 5 The problem of drug resistance

Although there are six anti-HIV PR inhibitors, and many other anti-reverse transcriptase drugs that are commercially available,
**Fig. 10**  FDA approved HIV-1 protease inhibitors.

**Fig. 11**  Noncleavable transition-state isostere developed for the synthesis of HIV PR inhibitors.
their effectiveness has been hampered by the emergence of drug-resistant and cross-resistant mutants, rendering AIDS with no definitive cure. The high rate of replication of the virus ($10^8$–$10^9$ virions/day) and the high error rate of HIV reverse transcriptase (about 1 in 10 000 bases) stand behind the rapid mutation and the selection of drug-resistant viruses. To date, according to the Stanford HIV reverse transcriptase and protease sequence database, 1929 isolates from patients exhibiting drug-induced protease mutations have been reported. Understanding the mechanism and structural basis of resistance in the light of the three-dimensional atomic structure of this enzyme has been reviewed.

Currently, there is no general strategy to tackle the problem of drug resistance. Modeling studies in our lab have shown that mutations that lead to drug resistance affect the hydrophobic binding site interacting with $P_3/P_3'$ and $P_1/P_1'$ groups of the inhibitors by reducing the $S_3$ binding region. We have also found that at least six mutated residues in HIV-1 PR causing drug resistance (K20I, V32I, I50V, N88D, L90M, Q92K) are found in the structurally aligned native residues of FIV PR (Feline immunodeficiency virus protease), which is also a $C_2$-symmetric homodimeric aspartic protease (Fig. 13) and has an identical mechanism of catalysis. Comparison of the structures of FIV and HIV-1 PRs and drug-resistant mutants of HIV PR further reveals a similarity between FIV PR and drug-resistant HIV PR. The FIV PR active site is more extended and contains a smaller $S_3$ subsite, a characteristic of many drug-resistant HIV proteases.

The above observations have suggested that FIV PR may serve as a model for drug resistance. Indeed, in a study to probe the specificity of $P_3$–$S_3$ interaction, it was found that inhibitors without a $P_3$ group or a small $P_3$ group not only exhibit a strong inhibition against the wild-type HIV protease, but also are effective against FIV protease and several drug-resistant mutants of HIV protease. One such inhibitor is TL-3 (Fig. 13), which was active in cell culture with IC$_{50}$ of 0.1, 0.4 and 0.9 µM for the wild-type, the I84V and F84V mutants, respectively, and no resistance was observed over a 6-month period.

### 6 Conclusion and perspectives

A few questions related to the detailed mechanism of HIV PR remain to be answered: the protonation state of the two aspartic groups, the location of the acidic proton, the role of the two flaps in the catalytic mechanism. However, it is clear from all the previous studies that many researchers agree about the similarity of the mechanism to that of the other aspartic proteases: the involvement of the water molecule as nucleophile, the general role of the two active site aspartyl residues. Although the HIV PR mechanism has not been completely characterized yet, these studies have driven several academic and industrial groups to develop a huge number of inhibitors based on the transition state concept, six of which are commercially available for the treatment of AIDS.

Many successes have been achieved in the war against AIDS, yet AIDS still has no definitive cure. All commercially available drugs and their combination in what is called the highly active antiretroviral therapy (HAART) improve the quality of life of...
the infected person. The current therapy is very expensive (95% of people who are infected with HIV/AIDS live in developing countries) and has other limitations such as the necessity of high doses, side effects and, mainly, the drug resistance problem.

Current research directed towards development of new therapies to cure AIDS is going in several directions. For example, development of new protease inhibitors that are not cross resistant to current drugs, inhibition of HIV-1 entry, and targeting conserved HIV RNA sequences with small molecules. Perhaps the ultimate solution is to develop a vaccine against this devastating epidemic.

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